



Europäisches Patentamt
European Patent Office
Office européen des brevets



(11) Publication number: **0 157 327 B1**

(12)

EUROPEAN PATENT SPECIFICATION

- (45) Date of publication of patent specification: 12.06.91 (51) Int. Cl.⁵: **C07D 207/36, C07D 333/32, C07D 307/58**
- (21) Application number: 85103507.1
- (22) Date of filing: 25.03.85

(54) Novel compounds for detecting the presence of hydrolytic analytes in a test sample.

- (30) Priority: 06.04.84 US 597336
- (43) Date of publication of application: 09.10.85 Bulletin 85/41
- (45) Publication of the grant of the patent: 12.06.91 Bulletin 91/24
- (84) Designated Contracting States:
AT BE CH DE FR GB IT LI LU SE
- (56) References cited:
EP-A- 0 094 554
US-A- 4 278 763
- Momose et al, Chem. Pharm. Bull, 27,
1148(1979)

- (73) Proprietor: MILES INC.
1127 Myrtle Street
Elkhart Indiana 46514(US)
- (72) Inventor: Ward, Frederick E.
deceased(US)
Inventor: Yip, Kin Fai
51194 Creekhaven Dr.
Elkhart, IN 46514(US)
Inventor: Yip, Meitak Teresa
51194 Creekhaven Dr.
Elkhart, IN 46514(US)
Inventor: Corey, Paul F.
1529 Middlebury St.
Elkhart, IN 46516(US)
- (74) Representative: Adrian, Albert, Dr. et al
c/o BAYER AG Konzernverwaltung RP Paten-
tabteilung
W-5090 Leverkusen 1, Bayerwerk(DE)

EP 0 157 327 B1

Note: Within nine months from the publication of the mention of the grant of the European patent, any person may give notice to the European Patent Office of opposition to the European patent granted. Notice of opposition shall be filed in a written reasoned statement. It shall not be deemed to have been filed until the opposition fee has been paid (Art. 99(1) European patent convention).

Description

The present invention relates to novel compounds useful in assaying a test sample for the presence of certain analyte constituents. Such analytes include leukocytes, esterase, and protease. The detection of leukocytes in urine is especially important in medical diagnostics.

The presence of an abnormal level of leukocytes in a patient's urine is possibly indicative of such pathological conditions as kidney or urogenital tract infection or other dysfunction. Accordingly, accurate urinary leukocyte information can be an invaluable tool to the physician in diagnosis and treatment of such pathologies.

Traditionally, the medical profession has relied on visual determination techniques to count leukocyte population in urine sediment or uncentrifuged urine, a process requiring expensive equipment such as a centrifuge and microscope, as well as inordinate time expenditure on the part of the clinician. Moreover, the traditional techniques suffer from the inadequacy that only intact cells are determined. Leukocytes occurring in the urinary system are subject to conditions which can favor extensive cell lysis. For example, it is known that in urines of abnormally high pH, leukocyte half life can be as low as 60 minutes. Since lysed cells escape detection in visual examination techniques, erroneously low determinations and false negatives can result.

Of the two techniques of microscopic leukocyte analysis - urine sediment and non-centrifuged, homogenized urine - the former is clearly the most desirable. Although dependable results can inure to the latter, urine sediment observation is used in an overwhelming majority of instances. It requires that the urine sample be centrifuged and the sediment isolated and subjected to microscopic inspection. The analyst then counts the number of leukocytes appearing in the viewing field. However, this task is further complicated by the presence of other urinary components in the sediment such as epithelial cells and salt particles. The varying content of sediment constituents, coupled with other complicating factors including non-homogeneity of the sample and differing optical powers among microscope equipment, can lead to enormous errors in the ultimate determination.

It is thus apparent that a quick, facile method of leukocyte determination, one which would eliminate the need for time-consuming techniques, as well as cost-consuming equipment, and which would provide accurate responses to esterase, protease or leukocyte cells, whether the cells were intact or had been lysed, would indeed advance the state of the art by a quantum jump. The present invention provides such an advance. Moreover, it is based, not on the ability to see leukocytes, but on the enzymatic activity they exhibit, and therefore is substantially free of the inaccuracies described above.

There exists in the prior art a body of references which disclose the use of certain esters which, when cleaved by enzymatic activity, result in the formation of color or other detectable species. Thus, British Patent No. 1,128,371 discloses the use of indoxyl and thioindoxyl esters as useful chromogens in detecting hydrolytic enzymes in body fluids. The enzymes cleave the ester to generate free indoxyl, which subsequently oxidizes to form the dimeric product indigo, a readily observable blue dye. Such activity is said to be due to, among other enzymes, cholinesterase. This patent also teaches that, in addition to the indoxyl portion of the ester substrate, the acid radical is chosen with particular reference to the enzyme to be detected. For example, it is stated that the acid radical can be acetate, or laurate or stearate for detection of esterase or lipase, respectively. For detecting enzymes such as phosphatase or sulfatase the acyl radical can be inorganic. Thus, the British Patent can be held to teach the use of chromogenic esters as substrates for determining esterolytic enzymes, such esters comprising indoxyl or thioindoxyl as the alcoholic moiety of the ester, the acyl moiety being tailored to be conducive to the particular enzyme to be determined.

The effect of careful acyl radical selection is nowhere more clearly exemplified than in two references which demonstrate esterase specificity for esters in which the acyl radical comprises an N-protected amino acid or peptide. Thus Janoff, *et al.*, *Proc. Soc. Exper. Biol. Med.* 136:1045-1049 (1971) teaches that alanine esters are specific substrates for esterase obtained from human leukocytes. Specifically this reference teaches that an extract of human leukocyte granules is capable of hydrolyzing N-acetyl-L-alanyl-L-alanyl-L-alanine methyl ester. Moreover, L-alanine-p-nitrophenol ester was similarly hydrolyzed to yield the yellow p-nitrophenol colorform.

Similarly, Sweetman *et al.*, *Jour. Hist. Soc.*, 22:327-339 teaches the use of 1-naphthyl-N-acetyl-DL-alanine and 1-naphthyl butyrate to demonstrate the presence of esterase, as well as 1-naphthyl-N-acetyl-L-alanyl-L-alanine.

United States Patent No. 4,278,763, assigned to Boehringer Mannheim GmbH combines these teachings in arriving at the indoxyl or thioindoxyl esters of amino acids or peptides as still another example of a traditional chromogenic substrate for leukocytic esterase activity. Moreover the Boehringer patent teaches

the equivalence of proteases and esterase in their estolytic pendants.

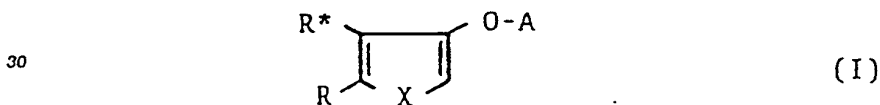
It is known that ester hydrolysis reactions can be activated by the presence of many nucleophilic agents, including many alcohols. Thus, the rate of hydrolysis of phenyl acetate and *p*-nitrophenyl acetate by esterase is increased 2.5 to 5.5 times upon addition of methanol and butanol. Greenzaid and Jencks, *Biochemistry*, 10(7), 12101227 (1971). Moreover, the effect increases with the length of the *n*-alkyl group. Wynne and Shalatin, *Eur. J. Biochem.*, 31, 554-560 (1972).

In particular, this activation effect of alcohols has been observed with esters of amino acids. *p*-Nitrophenyl *N*-acetyl-L-alaninate hydrolysis is activated (accelerated) by the presence of methanol. Fastrez and Fersht, *Biochemistry*, 12(11), 2025-2034 (1973). High molecular weight alcohols increase the rate of esterase-induced hydrolysis of *p*-nitrophenyl *t*-BOC-L-tyrosinate. Ashe and Zimmer, *Biochem. and Biophys. Res. Comm.*, 75(1), 194-199(1977). The disclosure of U.S. Patent No. 4,299,917 describes other known ester hydrolysis activators such as certain metal complexes, pyramine derivatives and imidazoles.

Also known is the use of certain diazonium salts to couple with phenols and pseudophenols to produce azo dyes. Martinet and Domier *Compt. Rend.*, 170, 592 (1920). Such a technique is used in an esterase analysis whereby indoxyl acetate is hydrolyzed via esterase to produce indoxyl, which is in turn coupled with a diazonium salt to form the corresponding azo dye. Holt and Hicks, *J. Cell Biol.* 29, 361-366 (1966); Gossrau, *Histochemistry*, 57, 323-342 (1978); West German *Offenlegungsschrift* No. 31 17 721, filed May 9, 1980.

The dye industry, an art which is clearly not analogous to medical diagnostic assays, but which nevertheless has long been a reservoir of experimental organic chemistry procedures, provides some guidance as to certain amino acids and their cyclization reactions. Thus, it is known to form pyrroles through the addition of glycine to salts of β -phenylglycidic acid, followed by treatment with hot acetic anhydride. Madelung and Obermann, *Ber.*, 63, 2870 (1930).

The present invention provides a new compound which is useful in determining the presence of leukocytes, esterase and/or protease in a test sample. The invention also relates to a method for preparing the compound. The compound is one having the structure



in which:

35 A is -COCH₃ or



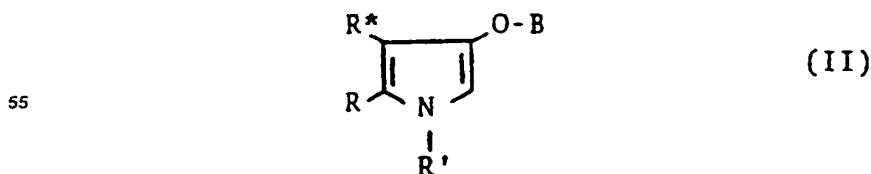
45 where Ts is tosyl;

R is a lower alkyl group having 1 to 6 carbon atoms, phenyl or chlorophenyl;

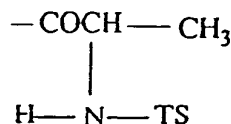
R* is H or a lower alkyl group having 1 to 6 carbon atoms; and

X is NR', in which R' is H or a lower alkyl group having 1 to 6 carbon atoms.

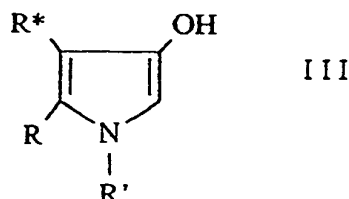
50 The presently claimed method relates to preparing the ester (I) wherein the structure is



in which B is $-\text{COCH}_3$ or

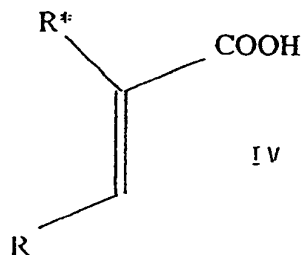


where Ts is tosyl; R is a lower alkyl group having 1 to 6 carbon atoms, phenyl or chlorophenyl. R' is H or a lower alkyl group having 1 to 6 carbon atoms; and R' is H or a lower alkyl group having 1 to 6 carbon atoms, the method comprising the sequential steps of
a) forming a 3-hydroxypyrrole having the structure



wherein R, R* and R' are as defined above, by the sequential steps of

1) reacting an aqueous mixture of a ketone, an alkali metal monopersulfate and a compound having the structure



wherein R and R* are as defined above, in the presence of a sufficient amount of alkali metal bicarbonate to maintain the mixture at a pH of at least 7 to form a first reaction mixture,

2) adding $\text{HOOC}-\text{CH}_2-\text{NHR}'$ to the first reaction mixture to form a second reaction mixture, R' being as defined above;

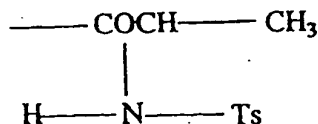
3) drying the second reaction mixture;

4) adding a symmetrical, lower alkyl carboxylic acid anhydride to the dried second reaction mixture in the presence of organic base to form a third reaction mixture;

5) hydrolyzing the resultant third reaction mixture to produce a reaction mixture containing a 3-hydroxypyrrole; and

6) isolating the 3-hydroxypyrrole;

b) adding an acid halide having the acyl group $-\text{COCH}_3$ or



where Ts is tosyl, to the 3-hydroxypyrrole in the presence of a carboxylic acid, and
c) isolating the resulting ester.

The following definitions are provided to clarify the scope of the present invention, and to enable its formulation and use.

The expression "lower alkyl", as used in the present disclosure, is an alkyl moiety containing about 1-6 carbon atoms. Included in the meaning of lower alkyl are methyl, ethyl, n-propyl, isopropyl, n-butyl, sec-butyl, tert-butyl and all isomers of pentyl and hexyl.

The alcoholic moiety of the ester is a pseudophenol in that it contains a heteroatom having a pair of electrons which can delocalize to produce aromaticity in concert with the double bonds in the ring.

Thus the heteroatom X is nitrogen, the atom can be unsubstituted i.e., NH, or it can be substituted with a lower alkyl. As a result of this aromatic character, the compound lends itself easily to coupling with a diazonium salt to form an azo dye, once the ester has become hydrolyzed.

It has been found that various substituents at the 5-position of the ring provide enhanced color formation and storage stability. Thus, R can be lower alkyl, phenyl, or p-chlorophenyl, provided that any such substituent group not interfere with the operation or functioning of the composition or test device in its capability to detect leukocyte cells, esterase or protease.

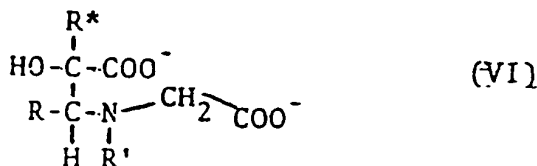
The acyl moiety A may be $-\text{COCH}_3$ or a N-tosyl-L-alanine residue.

The invention also includes a novel method of synthesizing the above compounds. The first step of the procedure comprises the formation of a 3-hydroxypyrrole having the structure (III). This is accomplished by a multistep synthesis beginning with a compound of structure (IV) as the starting material. This compound is reacted in the presence of acetone or other ketone to form an epoxide compound of the structure



The epoxidation is preferably achieved using an alkali metal monopersulfate such as KHSO_5 which is available commercially as Oxone® (DuPont Company). It has been found that pH control in this step ¹, ²is dramatically facilitated when the reaction is performed in the presence of excess alkali metal bicarbonate such as NaHCO_3 , i.e., at a pH of at least 7. Preferably the reaction mixture is kept at a temperature in the range of about 20 to 30 °C, although this range is not critical. Acidification of the resultant reaction mixture, followed by the addition of an organic solvent such as methylene chloride, serves to extract the epoxide into the organic phase.

Next the organic phase is washed with aqueous base. This step renders the epoxide water-soluble, thus transferring it to the aqueous phase. The aqueous phase is separated and a glycine added to produce a reaction mixture containing the double salt

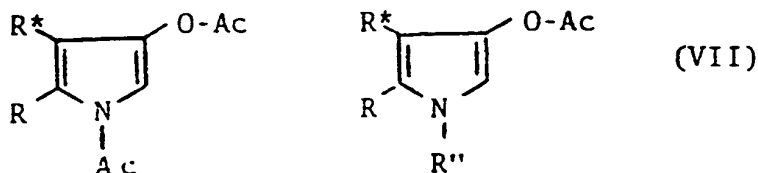


This step has been found to be best performed where the aqueous base has a pH range of about 10.5 to 12, with about 11.5 being the preferred pH. The temperature for this step is preferably elevated to about 100 °C, and a substantial amount of low boiling liquid removed (acetone and water). As the vapor temperature approaches 100 °C heating is maintained to continue reflux at or near 100 °C. Following cooling, washing with an organic solvent such as CH_2Cl_2 and evaporating to dryness (removal of residual water), (VI) is isolated from the residue by boiling in a solvent such as ethanol and allowing crystals to separate.

Next, compound (VI) is treated with a carboxylic acid anhydride to form

1. J.O. Edwards, et al, Photochem. Photobiol. 30, 63 (1979)

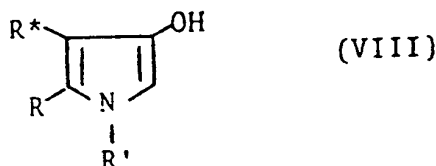
2. R. Curci, et al, J. Org. Chem. 45, 4758 (1980)



where R'' is lower alkyl or aryl. Ac is acetate in the case where the anhydride is acetic anhydride. Suitably this step is conducted initially at ambient temperature and ultimately heated to about 90 to about 140 °C. A key facet of this step is the inclusion of an organic base such as pyridine. Preferably Compound (VI) is suspended in the base, followed by the acid anhydride addition.

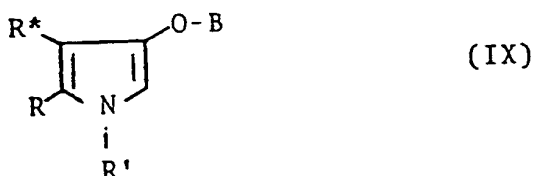
The thus-formed reaction mixture containing (VII) can then be evaporated to dryness to remove residual anhydride and solvent and the product (VII) recovered.

15 Treatment of (VII) under a hydrolyzing environment leads to the formation of



Preferred as a hydrolyzing environment is a mixture of an alcohol such as methanol and aqueous base, such as 2N NaOH. It is preferred to mix (VII), the alcohol and aqueous base at a low temperature, such as about -15 °C to about 10 °C. Compound (VIII) generally separates out of solution at this temperature and is isolated, such as by filtration, and purified.

The second step in the procedure is the esterification of (VIII) to form the product



in which B is as defined above.

(VIII) is added to an anhydrous solvent, such as tetrahydrofuran (THF), diethyl ether and the like, containing an organic base such as pyridine and an organic acid such as trifluoroacetic, oxalic, citric, acetic or other carboxylic acid. Preferred for this step are pyridine and trifluoroacetic acid. Next an acyl halide is added slowly. When the reaction is complete it is quenched with water, optionally containing a buffer such as citric acid and ethyl acetate, and the product (IX) recovered.

The following examples are provided further to assist the reader in making and using the present invention. Thus, preferred embodiments are described in experimental detail and analyzed as to their utility. The examples are illustrative only, and are in no way intended as limiting the scope of the invention described and claimed herein.

1 General Information

50 In the following experimental discussion abbreviations are used as indicated:

g = gram
 kg = kilogram
 L = liter
 55 mL = milliliter
 M = molar
 mM = millimolar
 N = normal

eq = equivalents
 mol = gram molecular formula (moles)
 mmol = gram molecular formula $\times 10^{-3}$ (millimoles)
 aq = aqueous
 5 hr = hour
 TLC = thin layer chromatography

Infrared (IR) spectra were obtained with a Perkin-Elmer Model 710B or 237 infrared spectrophotometer as solutions in CHCl_3 unless otherwise noted; the 1602 cm^{-1} band of polystyrene film was used as an external calibration standard. Signals are reported as cm^{-1} .

10 Proton magnetic resonance (^1H NMR) spectra were obtained at 89.55 MHz using a JEOL FX-90Q spectrometer or at 60 MHz using a Varian T-60 spectrometer; spectra were obtained in CDCl_3 solution unless otherwise noted. Chemical shifts are reported in parts per million downfield from the internal standard tetramethylsilane.

Carbon-13 magnetic resonance (^{13}C NMR) spectra were obtained at 22.5 MHz using a JEOL FX90Q
 15 spectrometer with Fourier transform and with full proton broad-band noise decoupling; spectra were obtained in CDCl_3 solution unless otherwise noted. Carbon shifts are reported in parts per million downfield from the internal standard tetramethylsilane.

Mass spectra (MS) were obtained on a Hewlett-Packard 5985A spectrometer operating in either an electron impact (EI) or fast atom bombardment (FAB) mode. High-resolution mass spectra were obtained on
 20 an AEI MS-902 spectrometer.

Organic reagents were obtained from Aldrich Chemical Company and were used without purification, unless otherwise noted. Inorganic reagents were ACS reagent grade from Fisher Scientific Company or other major vendor. Reaction solvents were ACS reagent grade; tetrahydrofuran (THF) was HPLC grade from J.T. Baker Chemical Company. Brine refers to a saturated aqueous sodium chloride solution.

25 Thin layer chromatograph (TLC) was performed using silica gel 60F-254 plates from E. Merck. Column chromatography was performed using E. Merck Silica Gel 60 (70-230 mesh). All melting points and boiling points reported are uncorrected.

2.1 Synthesis of 3-(*N*-tosyl-L-alaninyloxy)-5-phenylpyrrole (4)

30

The synthesis of (4) is illustrated in the following reaction sequence:

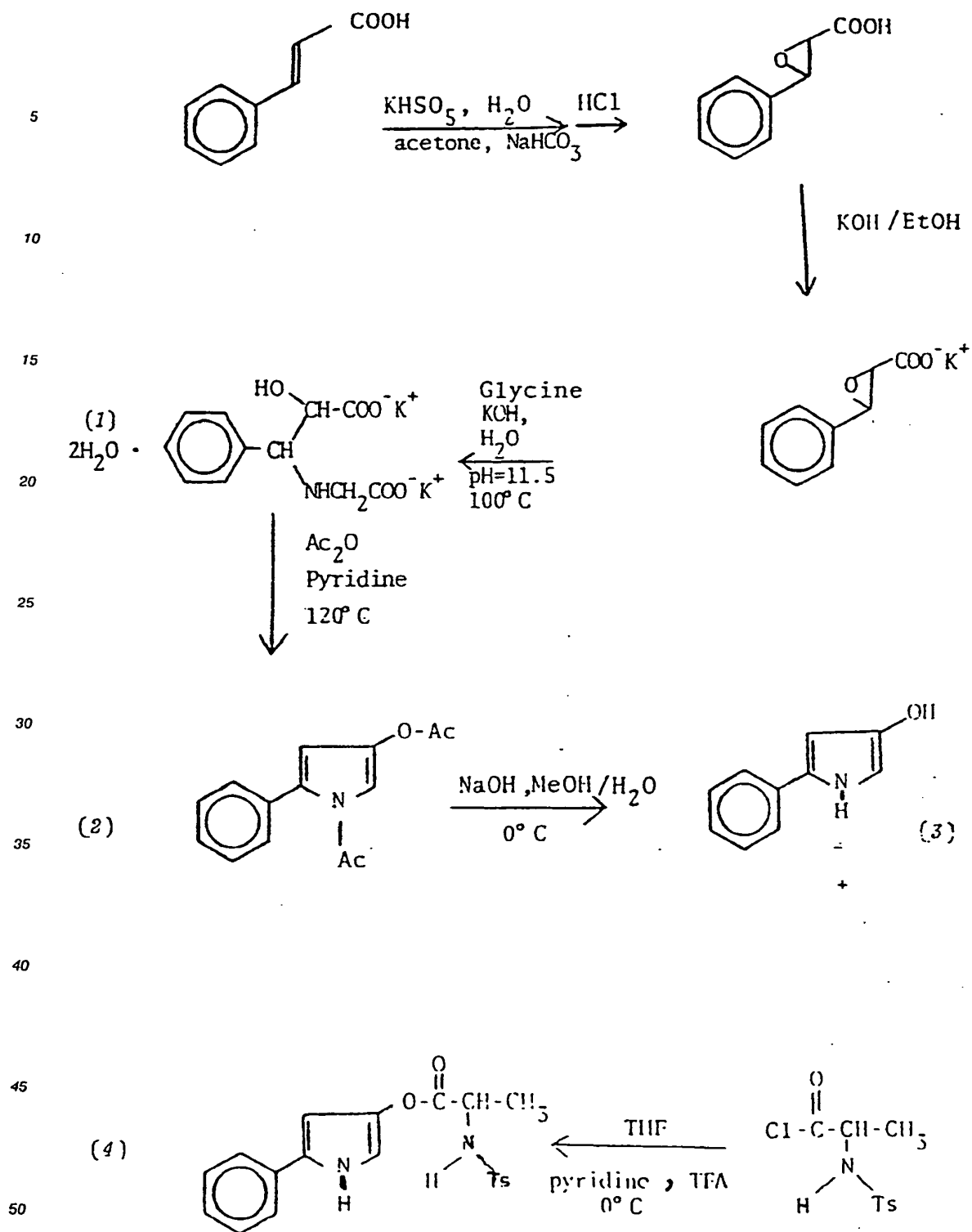
35

40

45

50

55

*N*-tosyl-L-alanine

L-alanine (100g; 1.11 moles) was dissolved in 2.25 L of 1 N sodium hydroxide (aq), cooled to 5° C and stirred while a solution of *p*-toluenesulfonyl chloride (218 g; 1.11 moles) in 450 mL of toluene was added slowly. The mixture was stirred at ambient temperature for 20 hr. The layers were separated and the chilled

aqueous layer acidified to pH 1 with concentrated hydrochloric acid. The white solid title compound was collected by filtration, washed with water and dried. Yield 178.5 g (66%) mp 134-5° C. IR (CHCl₃) cm⁻¹ 1726, 1340, 1165, 1095; ¹H NMR (DMSO-D₆) δ 1.20 (d, J=7,3H), 2.40 (s, 3H), 3.85 (p, J=8,1H), 6.4 (br d, 1H)(CO₂H), 7.41 (d, J_{AB} = 8, 2H) and 7.75 (d, J_{AB} = 8,2H) [center of pattern: 7.58; ΔV_{AB} = 20.49Hz], 8.03 (br d, J=8,1H)(NH).

N-tosyl-L-alaninyl chloride

Method A

A mixture of *N*-tosyl-L-alanine (12.4 g; 0.05 mol) and thionyl chloride (25 mL) was heated for 90 minutes at 55° C, and then concentrated on the rotary evaporator at 40° C. The red solid residue was dissolved in 200 mL of boiling CCl₄, decolorized with 20 g of oven dried Norit® 211 (American Norit Co., Inc.), filtered and chilled. The cream colored solid title product was collected by filtration, washed with hexane and dried. Yield 8.48 g (65%) with mp 101-101.5° C IR (CHCl₃) cm⁻¹ 3360, 3260, 3025, 1775, 1605, 1350, 1170, 910; ¹H NMR (CDCl₃) δ 1.48 (d, J=7,3H), 2.43 (s, 3H), 4.33 (p, J=8,1H), 5.93 (br d, J=8,1H)(NH), 7.31 (d, J_{AB} = 8, 2H) and 7.76 (d, J_{AB} = 8,2H) [center of pattern: 7.53; ΔV_{AB} = 26.83Hz].

Anal. calcd. for C₁₀H₁₂ClNO₃S: C, 45.89; H, 4.62; N, 5.35.

Found: C, 46.63; H, 4.90; N, 5.19.

Method B

A mixture of *N*-tosyl-L-alanine (3.1 g; 13 mmol) and thionyl chloride (6 mL) was heated for 90 min at 50° C, then diluted with 50 mL of dry hexane. The mixture was stirred rapidly, chilled and the solid product filtered. Yield 3.15 g (93%) mp 99-100° C. The IR spectrum was identical to that of the recrystallized material prepared by Method A.

2-Hydroxy-3(carboxymethylamino)-hydrocinnamic acid Dipotassium salt dihydrate (1)

A stirred slurry of 1.0 kg of *trans*-cinnamic acid (6.75 mol) in 4.5 L acetone was treated first with NaHCO₃ (2.47 kg; 29.4 mol; 4.36 eq) then carefully with water (4.5 L). The resulting thick mixture was treated dropwise, over 1.5-2.0 hr, with a solution of OXONE monopersulfate compound (3.78 kg; contains 1.825 eq of KHSO₅) in 0.4 mM aqueous disodium ethylenediamine tetraacetic acid (EDTA) (14.5 L; prepared by dissolving 2.17 g disodium EDTA dihydrate in 14.5 L distilled water). During this addition the reaction temperature was maintained at 24-27° C using a water bath; the reaction pH was noted to be about 7.4. After the addition was completed the mixture was stirred an additional 0.5 hr then cooled to about 10° C. The reaction was acidified with concentrated HCl (~ 1.2 L) to pH = 2, while maintaining the temperature at around 10° C, and then treated with CH₂Cl₂ (5.05 L) and stirred vigorously for 10 minutes.

After allowing the mixture to settle, the aqueous layer was decanted and set aside and the organic layer, which contained insoluble salts, was filtered through paper with suction. The filtered solids were washed with CH₂Cl₂ (1.9 L) and the aqueous layer extracted with this filtrate. The filtered solids were again washed with CH₂Cl₂ (3.15 L) and the aqueous layer extracted with this filtrate. The combined CH₂Cl₂ layers were extracted with a solution of KOH (593.3 g) in water (6.31 L). Gentle heating to about 40° C is often required to dissolve a solid which may separate during the base extraction. The CH₂Cl₂ layer was then extracted with a solution of KOH (99 g) in water (1.5 L) and the combined base extracts treated with glycine (481.7 g; 6.416 mol; 0.95 eq); the organic layer was discarded.

The aqueous solution pH was adjusted to 11.5 with 25% aqueous KOH then heated to boiling. Approximately 900 mL of low boiling liquid (acetone and water) was distilled off until the vapor temperature reached 99° C, following which, the mixture was refluxed for 2 hours. After cooling, the reaction mixture was extracted with CH₂Cl₂ (3.15 L), the CH₂Cl₂ phase discarded and the aqueous phase evaporated to dryness under reduced pressure with a 70° C bath. The residue was boiled in 95% ethanol (EtOH) (8.83 L) for 30 minutes, then allowed to cool slowly with stirring, whereupon the product separated as fine crystals. These were filtered, washed with fresh 95% EtOH (1.26 L) and dried in a 50-60° C oven to give the title compound (1.77 kg; 74.6%) as white crystals with mp = 120-2° C (uncorrected).

IR (KBr) cm⁻¹ 3420 (br.), 1590 (br.), 1410, 1130, 710; ¹H NMR (D₂O-TSP) δ 3.1 (s, 2H), 3.89 (d, J_{AB} = 4,1H) and 4.52 (d, J_{AB} = 4,1H) (center of pattern: 4.21; ΔV_{AB} = 18.83 Hz.), 4.68 (s, 6H, exchangeable

protons), 7.4 (s, 5H); TLC Rf = 0.59 (EtOH:1M triethylammonium bicarbonate, 7:3)

Anal. Calcd. for $C_{11}H_{15}NO_7K_2$: C, 37.59; H, 4.30; N, 3.99
Found: C, 37.22; H, 4.24; N, 3.96

N-acetyl-3-acetoxy-5-phenylpyrrole (2)

A suspension of 2-hydroxy-3-(carboxymethylamino)-hydrocinnamic acid dipotassium salt dihydrate (1) (1.0 kg; 2.84 mol) in pyridine (3.0 L) was treated with acetic anhydride (4.0 L) at ambient temperature under an inert gas atmosphere. A mild exothermic reaction ensued and the reaction temperature rose exponentially to 60-70 °C during a period of 1.5-2.5 hours. Once the reaction began to cool the mixture was heated to 120-123 °C for 15 minutes, then allowed to cool to ambient temperature over 1 hour, during which time pyridinium acetate separated as crystals. The mixture was filtered through paper with suction and the salts washed with ethyl acetate (EtOAc) until colorless; the filtrate was evaporated to dryness under vacuum.

The dark red residue was dissolved in EtOAc (3.0 L) washed three times with water (1.0 L) each, dried over $MgSO_4$ and treated with Darco®-G60 (ICI Americas, Inc.) (300 g). After stirring for 30 minutes the mixture was filtered through Celite® (Johns-Mannville) and evaporated to dryness under vacuum to give a reddish-orange oil. This oil was dissolved in warm 2-propanol (1.2 L), then allowed to cool slowly to ambient temperature overnight, whereupon a solid separates. The crystalline product was filtered, washed with 50% aqueous 2-propanol and dried to give the title compound (417 g; 60%) with mp = 58-60 °C (uncorrected). A portion was taken up in diethyl ether (Et_2O), treated with Norit 211, filtered and concentrated under reduced pressure; on standing at 0 °C colorless tiny needles separated. These were filtered, washed with Et_2O :Hexane (1:1) and vacuum dried to give the analytical sample with mp = 60-62.5 °C (uncorrected).

IR ($CHCl_3$) cm^{-1} 3020, 1760, 1730, 1595, 1378, 1320, 1220 (br.), 1030, 960, 903; 1H NMR ($CDCl_3$) δ 2.23 (s, 3H), 2.27 (s, 3H), 6.18 (d, J=2, 1H), 7.35 (s, 5H), 7.42 (d, J=2, 1H); TLC Rf = 0.56 (toluene:dioxane, 4:1).

Anal. Calcd. for $C_{14}H_{13}NO_3$: C, 69.12; H, 5.38; N, 5.76
Found: C, 68.88; H, 5.25; N, 5.53

3-Hydroxy-5-phenylpyrrole (3)

A finely divided portion of *N*-acetyl-3-acetoxy-5-phenylpyrrole (2) (36.8 g; 0.15 mol) was freed of oxygen by stirring in a flowing argon stream for 10 minutes, then suspended in deoxygenated methanol (MeOH) (379 mL), cooled to -6 °C (in a -15 °C methanol (MeOH)/dry-ice bath) under an inert gas atmosphere and rapidly treated with an ice cold deoxygenated solution of 2N NaOH (300 mL). The reaction temperature rose immediately upon addition of base to 18 °C, and after ~3 minutes the reaction mixture became homogeneous. As the reaction mixture cooled, compound 3 separated as fine crystals. After 15 minutes a solution of cold deoxygenated 2M citric acid (150 mL) was added, the resulting mixture was stirred for 10 minutes, and then filtered. The solid was washed thoroughly with deoxygenated water (200 mL), taking care to minimize exposure of the product to air, then dried under vacuum overnight to yield the title compound (22.3 g; 93.6%) as light pink tiny needles.

IR (KBr) cm^{-1} 3400, 3110, 2900, 1600, 1580, 1555, 1480, 1268, 1180, 742, 640; 1H NMR ($DMSO-d_6$) δ 6.1 (m, 1H), 6.3 (m, 1H), 7.0-7.7 (m, 5H), 8.0 (s, 1H), 10.4 (br s, 1H); TLC Rf = 0.20-0.28 (EtOH: $CHCl_3$, 1:9).

Anal. Calcd. for $C_{10}H_9NO$: C, 75.45; H, 5.70; N, 8.80
Found: C, 75.30; H, 5.69; N, 8.67

3-(*N*-tosyl-L-alanyloxy)-5-phenylpyrrole (4)

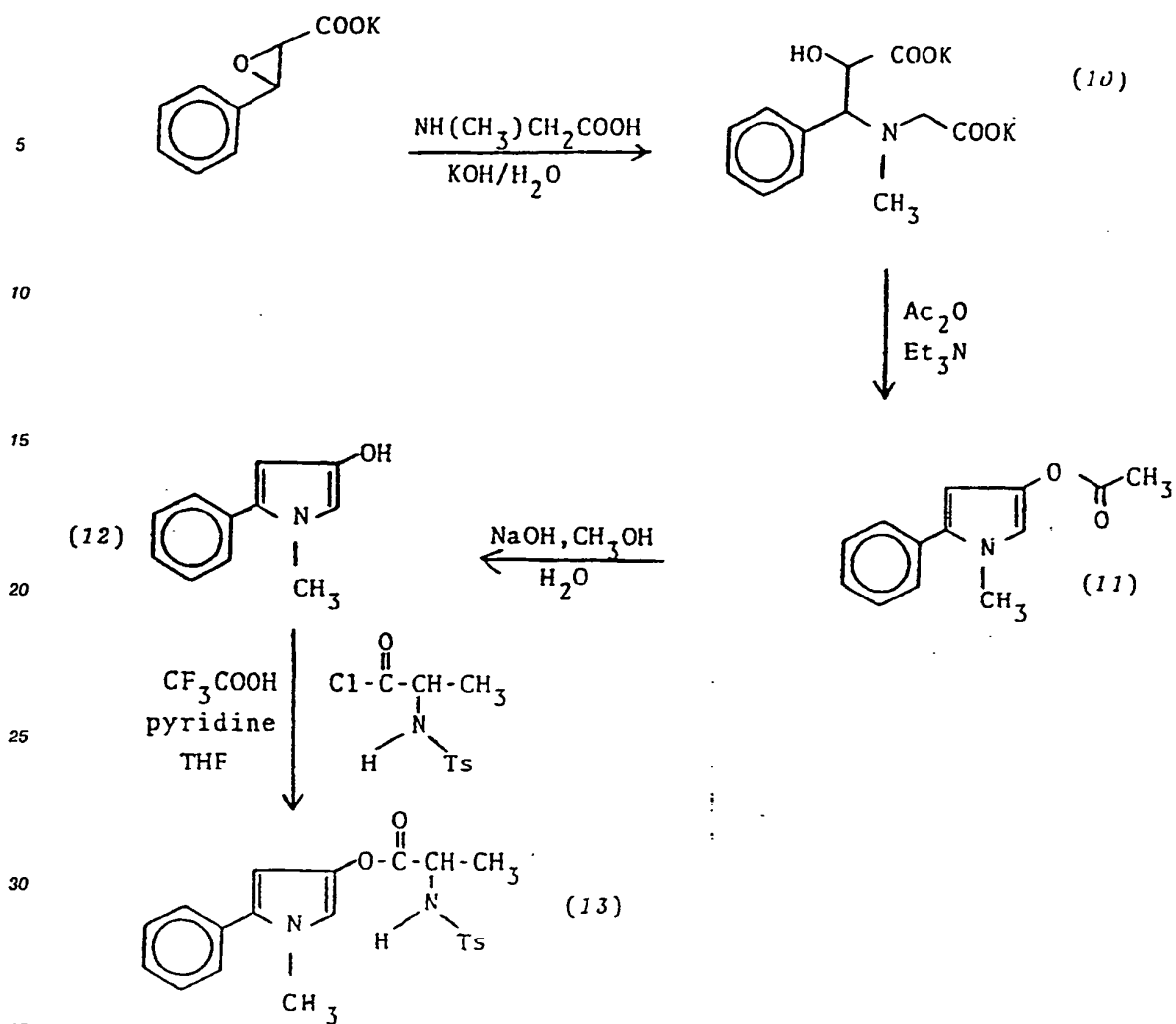
A solution of anhydrous tetrahydrofuran (THF, 450 mL), pyridine (43.8 mL; 0.542 mol; 1.2 eq) and trifluoroacetic acid (85.0 mL; 1.10 mol; 2.4 eq), maintained at 0 °C under an inert gas atmosphere, was treated in one portion with 3-hydroxy-5-phenylpyrrole (3) (71.5 g; 0.45 mol; 1.0 eq) followed immediately by the dropwise addition, over 5-10 minutes of a solution of freshly prepared N-tosyl-L-alaninyl chloride (141.0 g; 0.54 mol; 1.2 eq) in anhydrous THF (450 mL). The resulting mixture was stirred for 15 minutes at 0 °C. The reaction was then quenched by addition of a solution of 1.0 M aqueous citric acid (315 mL) and EtOAc (1.35 L). After brief mixing the phases were separated and the organic layer washed with a solution of aqueous NaCl (360 mL; 0.18 g NaCl per mL of water). The organic layer was next extracted twice with a solution of 5% aqueous NaHCO₃ (1.35 L each), and then washed with another portion of aqueous NaCl (360 mL; 0.18 g NaCl per mL of water). The reddish brown organic layer was stirred at ambient temperature for 15 minutes with MgSO₄ (101 g) and Darco-G60 (143 g), then filtered through Celite and evaporated to dryness under vacuum from a 37 °C bath to give (4) as a pinkish-white solid. The crude product was ground to a powder and dissolved in warm (50 °C) THF (250 mL), stirred vigorously and diluted with *n*-hexane (250 mL). The stirring was continued for 1 hour at ambient temperature as the product crystallized. The solid was filtered, washed with toluene (about 1 L) until the filtrate was colorless, then dried overnight to yield the title compound (112 g; 65%) as a white powder with mp = 154.5-155 °C.

IR (KCl) cm⁻¹ 3350, 3325, 1760, 1508, 1320, 1155, 770; ¹H NMR (DMSO-d₆) δ 1.33 (d, J = 7, 3H), 2.36 (s, 3H), 4.13 (p, J = 8, 1H), 6.25 (m, 1H), 6.73 (m, 1H), 7.05-7.50 (m, 5H), 7.5-7.85 (m, 4H), 8.42 (d, J = 8, 1H), 11.18 (br s, 1H); ¹³C NMR (DMSO-d₆) δ 18.335, 21.001, 51.370, 98.061, 108.336, 123.423, 126.024, 126.610, 128.560, 128.756, 129.601, 132.397, 137.600, 138.380, 142.737, 169.919; [α]_D = -70° (c = 1.11, MeOH); TLC R_f = 0.45 (EtOAc:hexane, 1:1); TLC R_f = 0.40 (toluene:dioxane, 4:1).

Anal. Calcd. for C₂₀H₂₀N₂O₄S: C, 62.48; H, 5.24; N, 7.29
Found: C, 62.62; H, 5.27; N, 7.30

2.3 Synthesis of 3-(N-tosyl-L-alaninyloxy)-1-methyl-5-phenylpyrrole (13)

A series of experiments was conducted to prepare the captioned ester corresponding to compound (II) in which A is N-tosyl-L-alaninyl, R is phenyl, R' is H, X is NR' and R' is CH₃. The reaction sequence is as follows:



2-Hydroxy-3-(N-methylcarboxymethylamino)-hydrocinnamic acid dipotassium salt (10)

A mixture of β -phenylglycidic acid potassium salt (30 g; 0.15 mole), N-methylglycine (13.2 g; 0.15 mole), distilled water (119 ml) and KOH solution (9N; 22.3 ml) was heated to reflux for 3 hours to give a light yellow solution. The reaction mixture was evaporated to dryness under reduced pressure at 70°C. The residue was then crystallized from 95% EtOH (100 ml) to give a white solid which, after drying overnight under reduced pressure at 110°C, yielded 30.8 g of white solid (10) (yield 63%).

IR (KCl) cm^{-1} 3360 (br.), 1580, 1405, 705; ^1H NMR (CD_3OD) δ 2.30 (s, 3H), 2.98 (s, 2H), 3.70 (d, $J=3$ Hz, 1H), 4.48 (d, $J=3$ Hz, 1H), 4.92 (s, 1H), 7.40 (s, 5H); TLC R_f = 0.51 (EtOH:1M triethylammonium bicarbonate, 7:3). (Product had no melting point less than 270°C).

3-Acetoxy-1-methyl-5-phenylpyrrole (11)

A mixture of 2-hydroxy-3-(N-methylcarboxymethylamino)-hydrocinnamic acid dipotassium salt (10) (15.2 g, 46 mmole), acetic anhydride (173 ml) and triethylamine (308 ml) was heated at 90°C for 19 hrs. The reaction mixture, which became deep brown in color, was filtered and the solid washed with ether. The filtrate was evaporated under reduced pressure to give a deep brown residue, which was taken up in ether (300 ml) and water (200 ml). The layers were separated and the ether layer washed with another portion of water (200 ml). The ether solution was then dried over MgSO_4 , filtered and concentrated under reduced pressure to give 10.7 g of brown residue which after Kugelrohr distillation and crystallization from ether yielded 3.0 g of white crystals (11) (yield 30%) mp = 64-65°C.

IR (CHCl₃) cm⁻¹ 2990, 1750, 1570, 1518, 1482, 1375, 1240 (br.), 1024, 910, 700; ¹H NMR (CDCl₃) δ 2.20 (s, 3H), 3.58 (s, 3H), 6.10 (d, J = 2 Hz, 1H), 6.75 (d, J = 2 Hz, 1H), 7.35 (s, 5H); TLC R_f = 0.58 (hexane: EtOAc 7:3)

⁵ Anal. Calcd. for $C_{13}H_{13}NO_2$: C, 72.54; H, 6.10; N, 6.44;
Found: C, 72.57; H, 6.09; N, 6.51

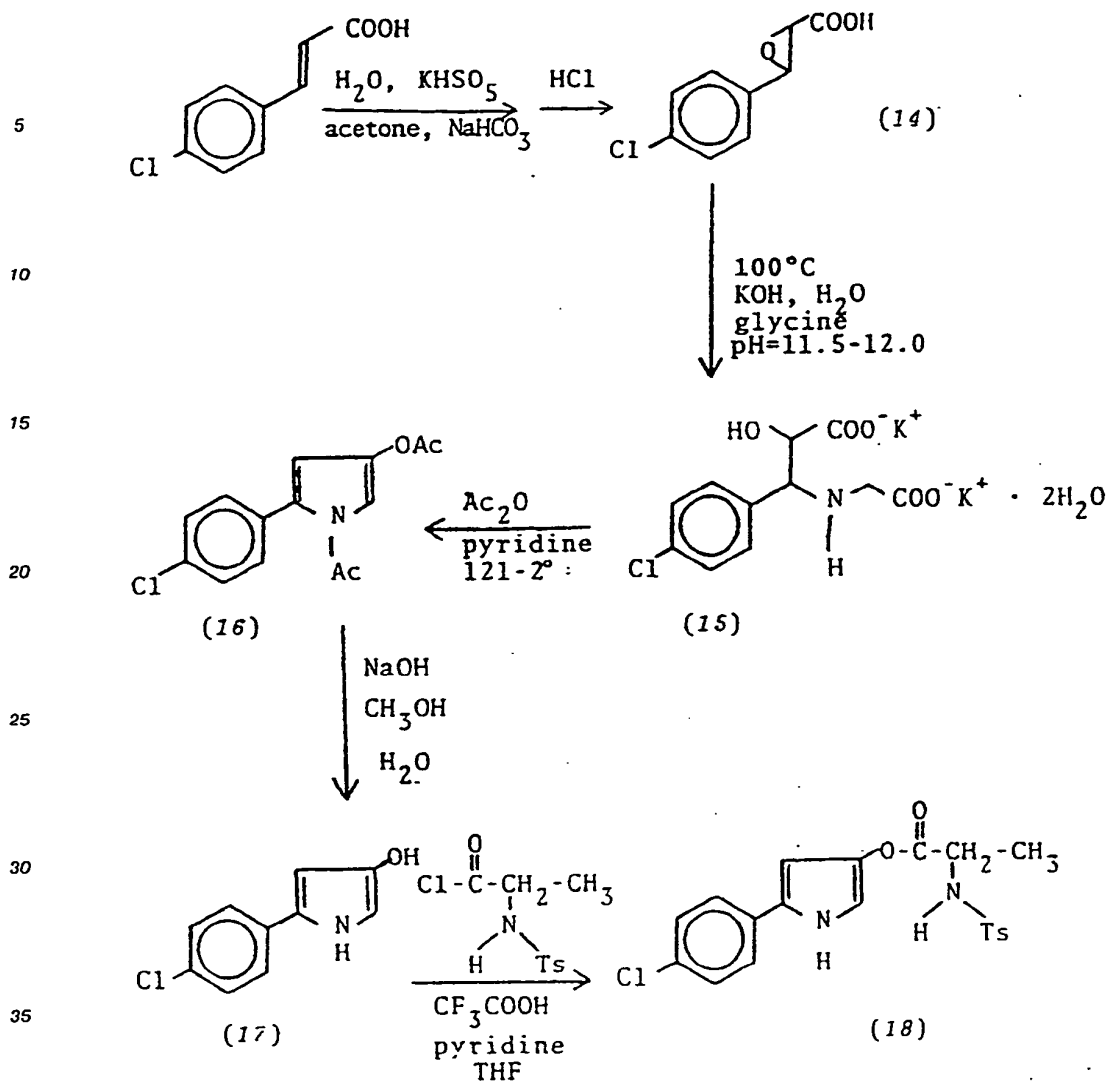
10 **3-(N-tosyl-L-alaninyloxy)-1-methyl-5-phenylpyrrole (13)**

To a mixture of deoxygenated methanol (15.5 ml) and 3-acetoxy-1-methyl-5-phenylpyrrole (**11**) (1.3 g, 6.2 mmole), under argon, was added deoxygenated NaOH (2N, 12.5 ml). The reaction mixture was stirred in an ice-bath for 15 minutes. Then deoxygenated citric acid (2M, 7 ml) was added and the resulting mixture was stirred in an ice bath for 8 minutes. The reaction mixture was concentrated under reduced pressure, then 20 ml of water was added and was extracted twice with ethylacetate (EtOAc) (50 ml). The EtOAc layers were combined, dried over MgSO₄, filtered and concentrated under reduced pressure to give 3-hydroxy-1-methyl-5-phenylpyrrole (**12**) as an orange residue. Under argon, a cold solution of anhydrous THF (12.4 ml), pyridine (0.6 ml, 7.4 mmole, 1.2 eq) and trifluoroacetic acid (1.2 ml, 15 mmole, 2.4 eq) was added to the orange residue, followed immediately by the addition of a solution of freshly prepared N-tosyl-L-alaninyl chloride (1.2 g, 7.4 mmole, 1.2 eq) in anhydrous THF (12.4 ml). The resulting mixture was stirred for one hr at 0 °C. Then the reaction was quenched by the addition of aqueous citric acid (1M, 5ml) and EtOAc (30 ml). After a brief mixing, the layers were separated and the organic layer was successively washed with saturated NaCl solution, twice with 5% NaHCO₃ solution and again with saturated NaCl solution. The EtOAc extract was then dried over MgSO₄, treated with Norit 211, filtered and concentrated under reduced pressure to give the crude product (**13**) as an orange residue. This was dissolved in hexane:EtOAc (1:1) (5 ml) and chromatographed on a column (SiO₂, 100 g) by elution with hexane:EtOAc (7:3) to give 1 g of (**13**) as a thick light orange oil. A portion of this crude product was further purified by semi-preparative HPLC (column, IBM silica, 1 cm x 25 cm; mobile phase, hexane:EtOAc 8:2; flow rate, 4.0 ml/min; pressure, 0.2 psi) to yield a honey color thick oil (**13**).

IR (film) cm^{-1} 3260, 2950, 1760, 1520, 1350, 1170, 770; ^1H NMR ($\text{DMSO}-d_6$) δ 1.28 (d, $J=7$ Hz, 3H), 2.36 (s, 3H), 3.58 (s, 3H), 5.85 (d, $J=2$ Hz, 1H), 6.15 (m, 1H), 6.74 (d, $J=2$ Hz, 1H), 7.30-7.80 (m, 9H), 8.37 (d, $J=8$ Hz, 1H); ^{13}C NMR ($\text{DMSO}-d_6$) ppm 8.205, 20.936, 34.917, 51.240, 100.598, 113.148, 126.544, 127.000, 128.105, 128.560, 129.601, 130.901, 132.202, 135.714, 138.315, 142.672, 169.724; TLC Rf = 0.52 (toluene:dioxane 4:1); High-resolution mass spectrum, $\text{C}_{21}\text{H}_{22}\text{N}_2\text{O}_4\text{S}$ requires m/e 398.1300, found m/e 398.1297.

6.2.4 Synthesis of 3-(N-tosyl-L-alaninyloxy)-5-(p-chlorophenyl)pyrrole (18)

A series of experiments was conducted to prepare the captioned ester compound corresponding to compound (I) in which A is *N*-tosyl-*L*-alaninyl, R is *p*-chlorophenyl, R' is H, X is NR' and R' is H. The reaction sequence is as follows:



40 *trans*- β -(*p*-Chlorophenyl)glycidic acid (14)

To a stirred slurry of *p*-chlorocinnamic acid (68.5 g; 0.375 mol) in 260 mL of acetone was added NaHCO_3 (137 g; 1.63 mol), followed by slow addition of 260 mL of water. To this mixture was added, over 2.5 hours at $22-27^\circ\text{C}$, a mixture of OXONE (211 g; 0.343 mol), 120 mg of disodium EDTA and 805 mL of water. After five hours the mixture was acidified with 70 mL of cold 12 N HCl, to bring the pH down to about 2.5, and it then was extracted with 700 mL of ethyl acetate. The extract was washed with brine, dried with MgSO_4 , filtered, and the filtrate was evaporated to dryness under vacuum. The white solids were crystallized from ethyl acetate: mp $121-5^\circ\text{C}$ (72.2 g; 97% yield). ^1H NMR ($\text{CDCl}_3/\text{DMSO}-d_6$) δ 7.3 (m, 4H), 4.05 (d, $J=2$, 1H), and 3.4 (d, $J=2$, 1H).

Anal. Calcd. for $\text{C}_9\text{H}_7\text{ClO}_3$: C, 54.43; H, 3.55; Cl, 17.85
Found: C, 54.53; H, 3.80; Cl, 17.91

2-Hydroxy-3-(carboxymethylamino)-p-chlorohydrocinnamic acid dipotassium salt dihydrate (15)

To a solution of KOH (85%) (46.7 g; 0.709 mol) and 400 mL of water was added glycine (25.9 g; 0.345 mol) followed by *trans*- β -*p*-chlorophenylglycidic acid (14) (72.2 g; 0.3635 mol). This mixture was heated at 100 °C for two hours, cooled to room temperature and sufficient KOH added to raise the pH to 12. The turbid solution was extracted three times with ethyl acetate, which extract was then discarded; the clear aqueous solution (about 500 mL) was evaporated under vacuum to dryness using a 70 °C water bath. The solids were then dissolved in about 350 mL of hot ethanol, filtered, and the filtrate chilled in an ice bath for several hours. The crystallized solids were collected by filtration and washed with some cold ethanol: mp 93-5 °C with decarboxylation at 185 °C (57.2 g; 41%).

¹H NMR (D₂O-TSP) δ 7.4 (s, 4H), 4.4 (d, J=4, 1H), 4.05 (d, J=4, 1H), and 3.1 (s, 2H).

**Anal. Calcd. for C₁₁H₁₀ClNO₅K₂·2H₂O: C, 34.24; H, 3.66;
N, 3.63**

Found: C, 34.40; H, 4.03; N, 3.42

N-Acetyl-3-acetoxy-5-(p-chlorophenyl)pyrrole (16)

To the 2-hydroxy-3-(carboxymethylamino)-*p*-chlorohydrocinnamic acid dipotassium salt dihydrate (15) (10 g; 0.02591 mol) was added acetic anhydride (40 mL) and pyridine (30 mL). This mixture was gently heated to 35 °C at which point the solution exothermed to 67 °C then began to fall, whereupon heating was again resumed. The mixture was heated at 121-2 °C (internal temperature) for one hour then cooled. To the reaction mixture was added about 30 mL of ethyl acetate which precipitated most the pyridinium acetate salt; this salt was collected by filtration and washed with a small amount of ethyl acetate. The filtrate was then evaporated under vacuum to an oil and ice water added. The product was extracted with ether and the ether extracts were successively washed twice with cold dilute aqueous citric acid, cold water, three times with cold dilute aq. NaHCO₃, cold water and brine, followed by drying over MgSO₄ and filtering. The filtrate was treated with 10 g of Darco, stirred for 20 minutes and then filtered. The filtrate was evaporated under vacuum to an oil. To the oil was added 25 mL of 2-propanol. The resultant solution yielded, with chilling and scratching, pale yellow crystals: mp 69-71 °C (3.4 g; 47%); TLC R_f = 0.61 (toluene:dioxane, 95:5). An analytical sample was recrystallized from 2-propanol but no change in mp was observed.

IR (KCl) cm⁻¹ 1755 (C=O, ester) and 1730 (C=O, amide); ¹H NMR (CDCl₃) δ 7.4 (m, 5H), 6.2 (d, J=2, 1H), 2.4 (s, 3H) and 2.3 (s, 3H).

Anal. Calcd. for C₁₄H₁₂ClNO₃: C, 60.55; H, 4.36; N, 5.04

Found: C, 60.65; H, 4.55; N, 5.07

3-Hydroxy-5-(p-chlorophenyl)pyrrole (17)

A sample of N-acetyl-3-acetoxy-5-*p*-chlorophenyl-pyrrole (16) (2.8 g; 0.01 mol) was deoxygenated for ten minutes with a stream of N₂. The solids were then dissolved in deoxygenated methanol (30 mL) which was then chilled to -8 °C. At once was added a cold deoxygenated solution of NaOH (1.6 g; 0.04 mol) in 20 mL H₂O, which solution was then heated briefly to 15 °C and then immediately cooled to -5 °C; after 25 minutes the clear solution was treated with a cold deoxygenated solution of citric acid (4.2 g; 0.02 mol) in 15 mL H₂O. The temperature rose briefly to 5 °C. After 0.5 hr. stirring at -5 °C, the solids were collected by filtration and washed with cold deoxygenated H₂O. The pale green product was dried under vacuum at room temperature over P₂O₅ for several days (1.3 g; 68%); TLC R_f=0.19 (CHCl₃:EtOH, 9:1); IR (KCl) showed no evidence for C=O absorption.

Anal. Calcd. for $C_{10}H_8ClNO \cdot 1/6H_2O$: C, 61.08; H, 4.27;
N, 7.12

Found: C, 61.36; H, 4.44;
N, 6.85

10 *3-(N-tosyl-L-alaninyloxy)-5-(p-chlorophenyl)pyrrole (18)*

To N_2 deoxygenated THF (15 mL) was added pyridine (0.65 mL; 0.008 mol), trifluoroacetic acid (1.27 mL; 0.0164 mol), and 3-hydroxy-5-(*p*-chlorophenyl)pyrrole (17) (1.3 g; 0.0065 mol). The solution was chilled to 0° C to -4° C and a N_2 deoxygenated chilled (0 to -4° C) solution of N-tosyl-L-alaninyl chloride (2.1 g; 0.008 mol) in 15 mL of THF was added over 10 minutes. After maintaining the mixture at 0° C for one hour, a mixture of ice and 100 mL of 1 N citric acid was added. This mixture was extracted with ethyl acetate and the extract washed once with cold brine, twice with cold dilute $NaHCO_3$, and once with cold brine, following which, it was dried over $MgSO_4$ and filtered. The filtrate was treated with 2 g of Darco and stirred for ten minutes, filtered and the filtrate concentrated under vacuum to a reddish-brown oil. A second treatment with 1.3 g Darco afforded a light reddish oil. The oil was dissolved in toluene:cyclohexane (4:1) and placed in the refrigerator overnight. Light salmon crystals were obtained. (1.45 g; 53%); mp 113-5° C; TLC Rf=0.47 (Et_2O); IR (KCl) cm^{-1} 1740 (C=O, ester); 1H NMR ($CDCl_3$) δ 8.4 (br s, 1H), 7.8-7.2 (m, 8H), 6.7 (m, 1H), 6.2 (m, 1H), 5.5 (d, J=9, 1H), 4.2 (p, J=8, 1H), 2.4 (s, 3H), 1.4 (d, 3H); MS (EI, DIP) m/e 418 (M^+ , 2.3%) and 420 (M^+ , 0.8%).

Anal. Calcd. for $C_{20}H_{19}ClN_2O_4S$: C, 57.34; H, 4.57;
N, 6.69

Found: C, 57.53; H, 4.58; N, 6.67

6.3 Preparation and Use of Test Devices Containing the Compound

35 A series of experiments was conducted to prepare test devices containing the present invention in which the ester substrates of paragraph 7.1, *supra*, were tested for responsiveness to leukocytes in urine. The devices comprised a small square of filter paper containing the assay reagents, the paper mounted at one end of a polystyrene film strip. The filter paper was impregnated with buffer, the ester an accelerator and a diazonium salt coupling agent. Each of the devices tested was found to exhibit a positive test for
40 leukocytes in urine.

6.3.1 Test device in which the ester is 3-(N-tosyl-L-alaninyloxy)-5-phenylpyrrole (4)

A test device, sensitive to the presence of leukocytes in urine, was prepared. The device comprised a
45 small square of filter paper mounted at one end of an oblong strip of polystyrene film. The paper was impregnated with various ingredients including a chromogenic ester, an accelerator and a diazonium salt. A 2 inch wide strip of Eaton and Dickman #205 filter paper was immersed in an aqueous solution containing the following:

0.4 M borate-NaOH buffer pH=8.6

50 2.0% (w/v) polyvinylpyrrolidone K-30

0.2% (w/v) Bioterger AS-40

0.25 M NaCl

The paper was then dried for 7 minutes in an Overly Air Foil paper dryer at 175-200° F at an airflow pressure of 1 inch of water. Next, the dried paper was immersed in an acetone solution containing

55 2.0% (v/v) *n*-decanol

0.75 mM 2-methoxy-4-morpholinobenzene diazonium chloride

0.5 mM 3-(N-tosyl-L-alaninyloxy)-5-phenylpyrrole

Following this impregnation the paper was dried for 10 minutes in a ventilated Hotpack® oven at 130° F. An

off-white test paper was obtained.

A piece of the dried, impregnated paper was cut to a square measuring 0.2 inches on a side and mounted at one end of an axially oriented polystyrene strip measuring 4 inches by 0.2 inches. Mounting the paper to the strip was achieved using Double Stick® double faced adhesive (3M Company).

5

6.3.3 Test device in which the ester is 3-(N-tosyl-L-alaninyloxy)-1-methyl-5-phenyl pyrrole (13)

Test devices were prepared following the procedure in experiment 6.3.1 in which the ester indicator was 3-(N-tosyl-L-alaninyloxy)-1-methyl-5-phenylpyrrole and the coupling agent was 1-diazonaphthalene-4-sulfonate. The aqueous first dip solution contained:

10

- 0.4 M boric acid
- 2.0% (w/v) polyvinylpyrrolidone (K-30)
- 0.2% (v/v) Bioterg AS-40
- 0.25 M NaCl

15

Prior to impregnation of the filter paper, the solution was titrated with NaOH to a pH of 9.0.

The second dip solution in acetone contained:

- 0.75 mM 1-diazo-2-naphthol-4-sulfonate
- 1.3 mM 3-(N-tosyl-L-alaninyloxy)-1-methyl-5-phenylpyrrole
- 1.5% (v/v) dodecanol

20

Following impregnation in the aqueous first dip, the paper was dried for about 5 minutes at about 80 °C, and for about 5 minutes at 70 °C following impregnation in the acetonic second dip.

The dried paper was mounted as in experiment 3.1.

25

6.3.4 Test device in which the ester is 3-(N-tosyl-L-alaninyloxy)-5-(p-chlorophenyl) pyrrole (18)

Test devices were prepared as in Experiment 6.3.1 except that the acetone solution contained, in place of the phenylpyrrole, 1.3 mM 3-(N-tosyl-L-alaninyloxy)-5-(p-chlorophenyl)pyrrole.

6.4 Evaluation of the Test Device

30

The test devices prepared in the above experiments were subjected to evaluation of their ability to detect leukocytes present in urine.

Test samples were prepared from a normal human urine pool. One sample served as a blank and leukocytes isolated from freshly drawn blood were added to two additional urine samples to yield concentrations of 0, 10 and 75 leukocytes/ μ L, respectively.

35

Test devices were quickly immersed in and removed from a test sample. Two minutes later the devices were observed using a spectrophotometer to measure % reflectance at different wavelengths from 400-700 nm (nanometers).

The data show that all of the test devices demonstrated clearly discernable differences in light reflectance corresponding to different leukocyte levels in the test samples. The data are presented in the following table.

40

45

50

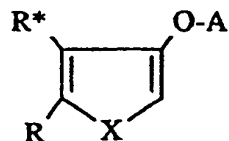
55

Experiment No.	Leukocyte Concentration (cells/ μ L)	% Reflectance at 555 nm
6.3.1*	0	-
	10-12	-
6.3.3	0	67
	10	64
	75	60
6.3.4	0	61
	10	51
	75	42

*Visual observation: *purple color formed at 10-12 cells/ μ L; blank gave no color change

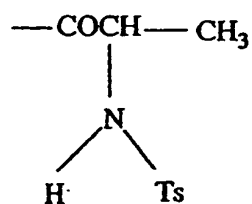
Claims

1. A compound having the structure



in which:

A is $-\text{COCH}_3$ or



where Ts is tosyl;

R is a lower alkyl group having 1 to 6 carbon atoms, phenyl or chlorophenyl;

R* is H or a lower alkyl group having 1 to 6 carbon atoms; and
X is NR', in which R' is H or a lower alkyl group having 1 to 6 carbon atoms.

2. A compound selected from the group 3-(N-tosyl-L-alaninyloxy)-5-phenyl pyrrole, 3-(N-tosyl-L-alaninyloxy)-1-methyl-5-phenyl pyrrole, 3-acetoxy-1-methyl-5-phenyl pyrrole and 3-(N-tosyl-L-alaninyloxy)-5-(p-chlorophenyl)pyrrole.

3. A method for preparing an ester having the structure



in which B is -COCH₃ or



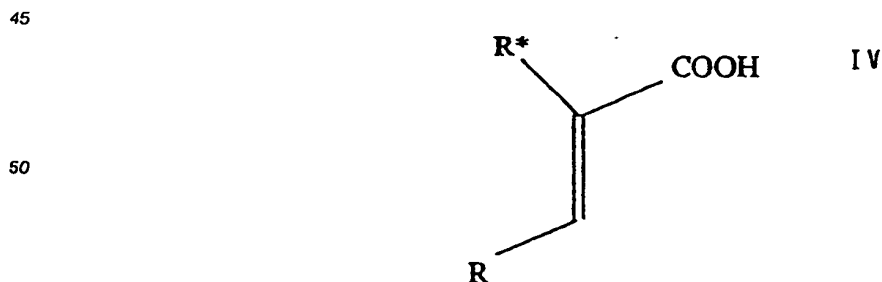
where Ts is tosyl; R is a lower alkyl group having 1 to 6 carbon atoms, phenyl or chlorophenyl. R* is H or a lower alkyl group having 1 to 6 carbon atoms; and R' is H or a lower alkyl group having 1 to 6 carbon atoms, the method comprising the sequential steps of

a) forming a 3-hydroxypyrrole having the structure



wherein R, R* and R' are as defined above, by the sequential steps of

1) reacting an aqueous mixture of a ketone, an alkali metal monopersulfate and a compound having the structure



wherein R and R* are as defined above, in the presence of a sufficient amount of alkali metal bicarbonate to maintain the mixture at a pH of at least 7 to form a first reaction mixture, 2) adding HOOC-CH₂-NHR' to the first reaction mixture to form a second reaction mixture, R'

- being as defined above;
 3) drying the second reaction mixture;
 4) adding a symmetrical, lower alkyl carboxylic acid anhydride to the dried second reaction mixture in the presence of organic base to form a third reaction mixture;
 5) hydrolyzing the resultant third reaction mixture to produce a reaction mixture containing a 3-hydroxypyrrole; and
 6) isolating the 3-hydroxypyrrole;
 b) adding an acid halide having the acyl group $-\text{COCH}_3$ or

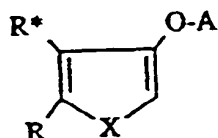


where Ts is tosyl, to the 3-hydroxypyrrole in the presence of a carboxylic acid, and
 c) isolating the resulting ester.

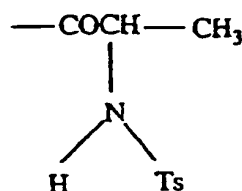
4. The method of claim 3 in which R is phenyl.
5. The method of claim 3 in which R is phenyl and R* is H.
6. The method of claim 3 in which the pH of the first reaction mixture is raised to a pH in the range of 9.6-12.
7. The method of claim 3 in which methanol, water and alkali metal hydroxide are employed to hydrolyze the third reaction mixture.
8. The method of claim 7 which comprises about 0.1 to about 4N alkali metal hydroxide, water and methanol.
9. The method of claim 3 in which the acid halide is acid chloride.
10. The method of claim 3 in which the carboxylic acid is trifluoroacetic acid.

Revendications

1. Composé répondant à la formule



dans laquelle
 A représente $-\text{COCH}_3$ ou



où Ts est le groupe tosyle ;

R est un groupe alkyle inférieur ayant 1 à 6 atomes de carbone, phényle ou chlorophényle ;

R* représente H ou un groupe alkyle inférieur ayant 1 à 6 atomes de carbone ; et

x représente NR', où R' représente H ou un groupe alkyle inférieur ayant 1 à 6 atomes de carbone.

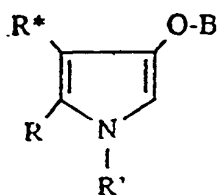
5

2. Composé choisi dans le groupe comprenant le 3-(N-tosyl-L-alaninyloxy)-5-phénylpyrrole, le 3-(N-tosyl-L-alaninyloxy)-1-méthyl-5-phénylpyrrole, le 3-acétoxy-1-méthyl-5-phénylpyrrole et le 3-(N-tosyl-L-alaninyloxy)-5-(p-chlorophényl)pyrrole.

10

3. Procédé de préparation d'un ester répondant à la formule

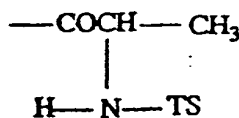
15



20

dans laquelle B est un groupe -COCH₃ ou

25



30

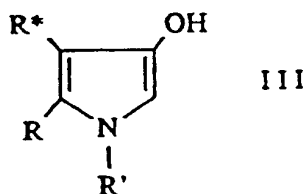
où Ts est un groupe tosyle ; R est un groupe alkyle inférieur ayant 1 à 6 atomes de carbone, phényle ou chlorophényle. R* représente H ou un groupe alkyle inférieur ayant 1 à 6 atomes de carbone et R' représente H ou un groupe alkyle inférieur ayant 1 à 6 atomes de carbone,

procédé comprenant les étapes successives consistant

35

- a) à former un 3-hydroxypyrrole de formule

40



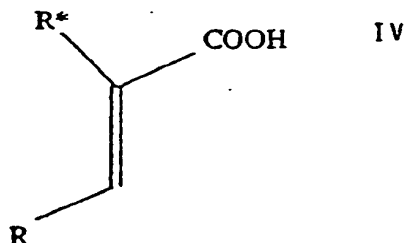
45

dans laquelle R, R* et R' sont tels que définis ci-dessus, par des étapes successives de

- 1) réaction d'un mélange aqueux d'une cétone, d'un monopersulfate de métal alcalin et d'un composé répondant à la formule

50

55



dans laquelle R et R* sont tels que définis ci-dessus, en présence d'une quantité suffisante d'un bicarbonate de métal alcalin pour maintenir le mélange à un pH au moins égal à 7 de manière à former un premier mélange réactionnel,

2) addition de $\text{HOOC-CH}_2\text{-NHR'}$ au premier mélange réactionnel pour former un second mélange réactionnel, R' ayant la définition donnée ci-dessus ;

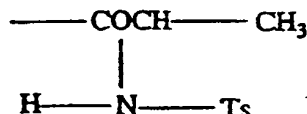
3) séchage du second mélange réactionnel ;

4) addition d'un anhydride d'acide alkylcarboxylique symétrique à groupe alkyle inférieur au second mélange réactionnel séché en présence d'une base organique pour former un troisième mélange réactionnel ;

5) hydrolyse du troisième mélange réactionnel résultant pour produire un mélange réactionnel contenant un 3-hydroxypyrrole et

6) isolement du 3-hydroxypyrrole ;

b) à ajouter un halogénure d'acide portant le groupe acyle -COCH_3 ou

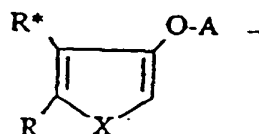


où Ts est un groupe tosyle, au 3-hydroxypyrrole en présence d'un acide carboxylique et
c) à isoler l'ester résultant.

- 35 4. procédé suivant la revendication 3, dans lequel R est un groupe phényle.
5. procédé suivant la revendication 3, dans lequel R est un groupe phényle et R* représente H.
6. Procédé suivant la revendication 3, dans lequel le pH du premier mélange réactionnel est élevé dans
40 une plage de 9,6 à 12.
7. Procédé suivant la revendication 3, dans lequel du méthanol, de l'eau et un hydroxyde de métal alcalin sont utilisés pour hydrolyser le troisième mélange réactionnel.
- 45 8. Procédé suivant la revendication 7, qui comprend un hydroxyde de métal alcalin d'environ 0,1 à environ 4N, de l'eau et du méthanol.
9. Procédé suivant la revendication 3, dans lequel l'halogénure d'acide est un chlorure d'acide.
- 50 10. Procédé suivant la revendication 3, dans lequel l'acide carboxylique est l'acide trifluoracétique.

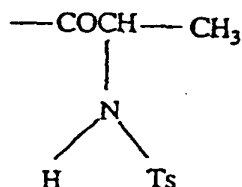
Ansprüche

- 55 1. Verbindung mit der Struktur



worin:

A -COCH₃ oder



worin Ts Tosyl ist,

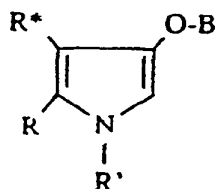
R eine Niedrigalkylgruppe mit 1 bis 6 Kohlenstoffatomen, Phenyl oder Chlorphenyl,

R* H oder eine Niedrigalkylgruppe mit 1 bis 6 Kohlenstoffatomen und

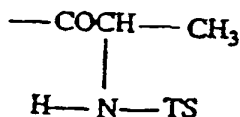
X NR' worin R' H oder eine Niedrigalkylgruppe mit 1 bis 6 Kohlenstoffatomen sind.

2. Verbindung, ausgewählt aus der Gruppe
 - 3-(N-Tosyl-L-alaninyloxy)-5-phenyl-pyrrol,
 - 3-(N-Tosyl-L-alaninyloxy)-1-methyl-5-phenylpyrrol,
 - 3-Acetoxy-1-methyl-5-phenylpyrrol und
 - 3-(N-Tosyl-L-alaninyloxy)-5-(p-chlorphenyl)pyrrol.

3. Verfahren zur Herstellung des Esters (I) mit der Struktur

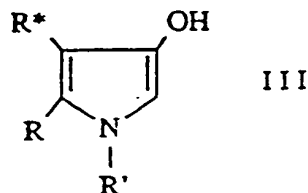


worin B -COCH₃ oder

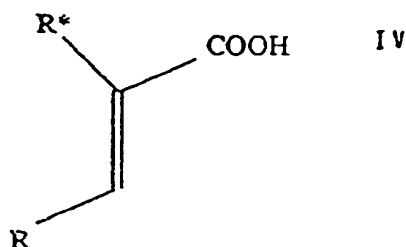


worin Ts Tosyl ist, R eine Niedrigalkylgruppe mit 1 bis 6 Kohlenstoffatomen, Phenyl oder Chlorphenyl, R* H oder eine Niedrigalkylgruppe mit 1 bis 6 Kohlenstoffatomen und R' H oder eine Niedrigalkylgruppe mit 1 bis 6 Kohlenstoffatomen sind, wobei das Verfahren die aufeinanderfolgenden Stufen umfaßt, wobei man

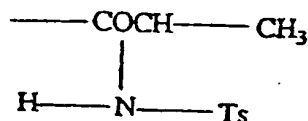
- a) ein 3-Hydroxypyrrol mit der Struktur bildet



10 worin R, R* und R' wie oben definiert sind, wobei man aufeinanderfolgend
1) eine wässrige Mischung eines Ketons, eines Alkalimetallmonopersulfats und einer Verbindung
mit der Struktur



25 worin R und R* wie oben definiert sind, in der Gegenwart einer zur Aufrechterhaltung eines pH in
der Mischung von mindestens 7 hinreichenden Menge an Alkalimetallbicarbonat reagieren läßt,
um eine erste Reaktionsmischung zu bilden.
2) $\text{HOOC-CH}_2\text{-NHR'}$ zur ersten Reaktionsmischung zufügt, um eine zweite Reaktionsmischung zu
bilden, worin R' wie oben definiert ist,
3) die zweite Reaktionsmischung trocknet,
30 4) ein symmetrisches Niedrigalkylcarboxylsäureanhydrid der getrockneten zweiten Reaktionsmischung
in der Gegenwart einer organischen Base zufügt, um eine dritte Reaktionsmischung zu
bilden,
5) die sich ergebende dritte Reaktionsmischung hydrolysiert, um eine Reaktionsmischung herzu-
stellen, enthaltend ein 3-Hydroxypyrrol, und
35 6) das 3-Hydroxypyrrol isoliert,
b) ein Säurehalogenid mit der Acylgruppe -COCH_3 oder



45 worin Ts Tosyl ist, dem 3-Hydroxypyrrol in der Gegenwart einer Carboxylsäure zufügt und
c) den sich ergebenden Ester isoliert.

4. Verfahren nach Anspruch 3,
worin R Phenyl ist.

50 5. Verfahren nach Anspruch 3,
worin R Phenyl und R* H sind.

6. Verfahren nach Anspruch 3,
wobei der pH der ersten Reaktionsmischung auf einen pH im Bereich von 9.6 bis 12 angehoben wird.

55 7. Verfahren gemäß Anspruch 3,
wobei Methanol, Wasser und Alkalimetallhydroxid eingesetzt werden, um die dritte Reaktionsmischung
zu hydrolysieren.

8. Verfahren gemäß Anspruch 7,
wobei man etwa 0,1 bis etwa 4N Alkalimetallhydroxid, Wasser und Methanol einsetzt sind.
9. Verfahren gemäß Anspruch 3,
wobei das Säurehalogenid ein Säurechlorid ist.
10. Verfahren gemäß Anspruch 3,
wobei die Carbonsäure Trifluoressigsäure ist.

10

15

20

25

30

35

40

45

50

55